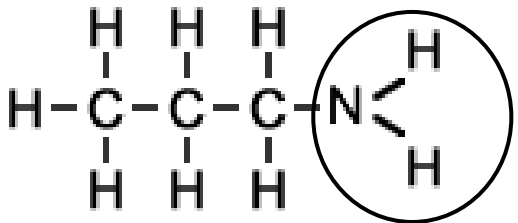
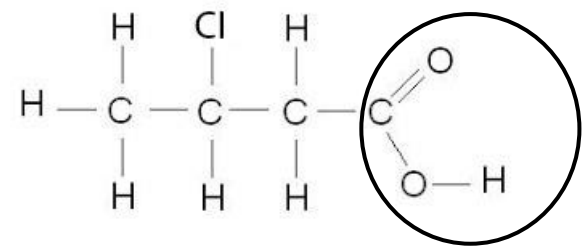
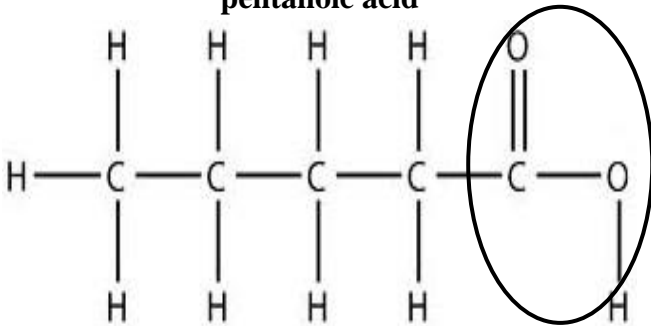
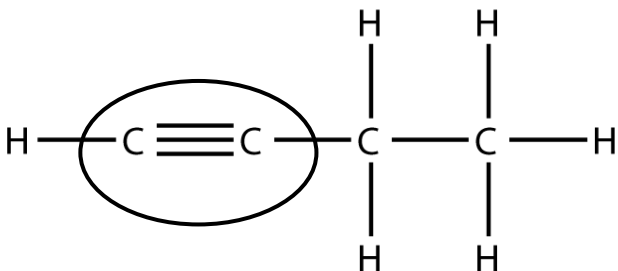
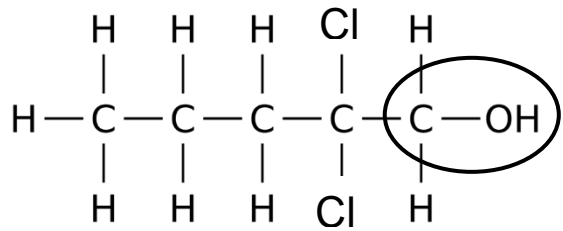
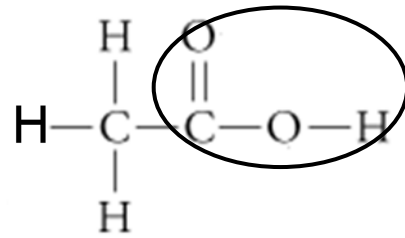
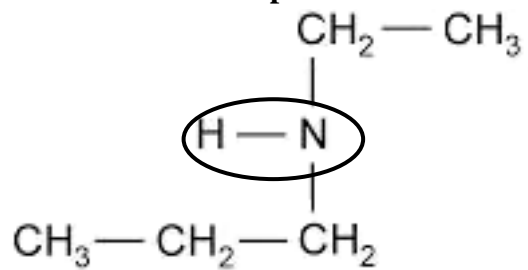
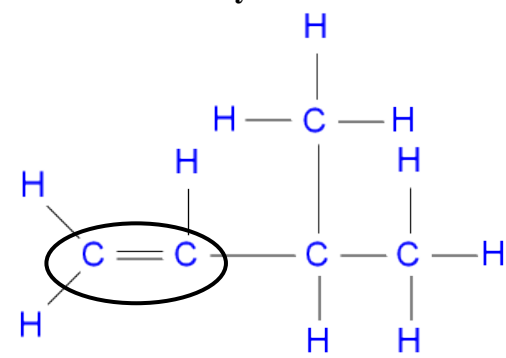
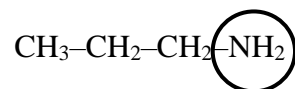


ANSWERS: Structural formula (aka displayed, expanded structural formula) practice

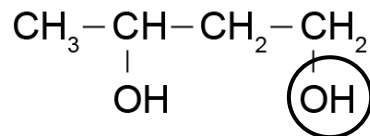
<p>3-methylpentanoic acid</p> <p>available in 2017</p>	<p>propan-1-amine</p>  <p>amine</p>	<p>2-chlorobutanoic acid</p>  <p>carboxylic acid</p>
<p>pentanoic acid</p>  <p>carboxylic acid</p>	<p>but-1-yne</p>  <p>alkyne</p>	<p>2,2-dichloropentan-1-ol</p>  <p>alcohol</p>
<p>ethanoic acid</p>  <p>carboxylic acid</p>	<p>2-aminopentane</p>  <p>amine</p>	<p>3-methylbut-1-ene</p>  <p>alkene</p>

1-aminopropane



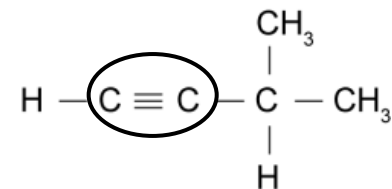
amine

butan-1,3-diol



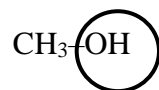
alcohol

3-methylbut-1-yne



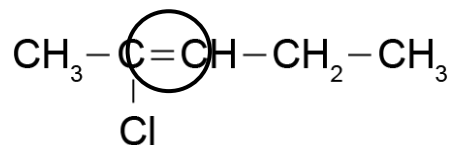
alkyne

methanol



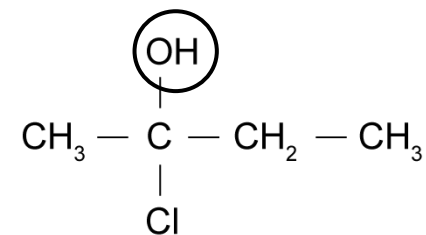
alcohol

2-chloropent-2-ene



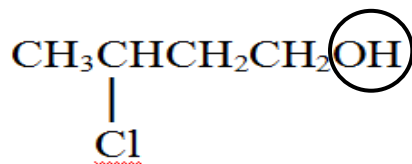
alkene

2-chlorobutan-2-ol



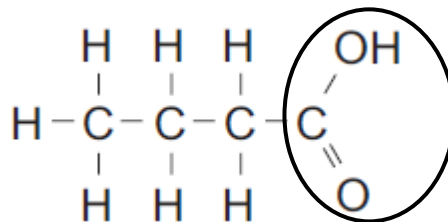
alcohol

3-chlorobutan-1-ol



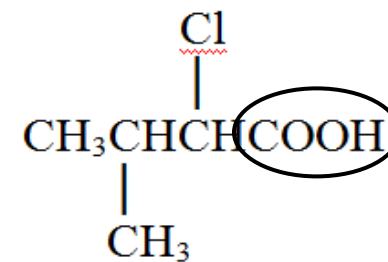
alcohol

C₄H₈O₂ butanoic acid



carboxylic acid

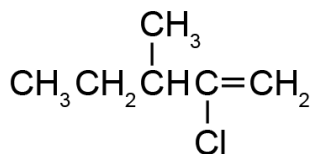
2-chloro-3-methylbutanoic acid



carboxylic acid

<p>A branched-chain secondary alcohol C₅H₁₁OH 3-methylbutan-2-ol</p> $\begin{array}{c} \text{OH} \\ \\ \text{H}_3\text{C}-\text{CH}-\text{CH}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$ <p>alcohol</p>	<p>A branched-chain tertiary alcohol C₅H₁₁OH 2-methylbutan-2-ol</p> $\begin{array}{c} \text{OH} \\ \\ \text{H}_3\text{C}-\text{C}-\text{CH}_2-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$ <p>alcohol</p>	<p>2-amino-3-methylbutane</p> $\begin{array}{c} \text{H}_3\text{C}-\text{CH}-\text{CH}-\text{CH}_3 \\ \quad \\ \text{NH}_2 \quad \text{CH}_3 \end{array}$ <p>amine</p>
<p>4-chlorobutanoic acid</p> $\begin{array}{c} \text{Cl} \\ \\ \text{H}_2\text{C}-\text{CH}_2-\text{CH}_2-\text{C}-\text{OH} \\ \\ \text{O} \end{array}$ <p>carboxylic acid</p>	<p>pentan-2-ol</p> $\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}-\text{CH}_3$ <p style="text-align: center;">OH</p> <p>alcohol</p>	<p>4-methylpent-2-yne</p> $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{C}\equiv\text{C}-\text{C}-\text{CH}_3 \\ \\ \text{H} \end{array}$ <p>alkyne</p>
<p>3-aminopentane</p> $\begin{array}{c} \text{NH}_2 \\ \\ \text{H}_3\text{C}-\text{CH}_2-\text{CH}-\text{CH}_2-\text{CH}_3 \end{array}$ <p>amine</p>	<p>2-chloropropan-1-ol</p> $\begin{array}{c} \text{OH} \quad \text{Cl} \\ \quad \\ \text{H}_2\text{C}-\text{CH}-\text{CH}_3 \end{array}$ <p>alcohol</p>	<p>2-amino-2,3-dimethylbutane</p> $\begin{array}{c} \text{CH}_3 \quad \text{CH}_3 \\ \quad \\ \text{CH}_3-\text{CH}-\text{C}-\text{CH}_3 \\ \\ \text{NH}_2 \end{array}$ <p>amine</p>

<p>pent-2-ene</p> $\text{CH}_3-\text{CH}=\text{CH}-\text{CH}_2-\text{CH}_3$ <p>alkene</p>	<p>3-chloropropan-1-ol</p> $\text{HO}-\text{CH}_2-\text{CH}_2-\text{CH}_2-\text{Cl}$ <p>alcohol</p>	<p>2-bromo-3-methylbutanoic acid</p> $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}-\text{CH}-\text{C}-\text{OH} \\ \quad \quad \\ \text{Br} \quad \quad \text{O} \end{array}$ <p>carboxylic acid</p>
<p>4-methylpent-2-ene</p> $\begin{array}{c} \text{CH}_3 \\ \\ \text{CH}_3-\text{CH}=\text{CH}-\text{CH}-\text{CH}_3 \\ \\ \text{CH}_3 \end{array}$ <p>alkene</p>	<p>butanoic acid</p> $\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C} \\ \quad \quad \quad // \quad \backslash \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{O} \quad \text{O}-\text{H} \end{array}$ <p>carboxylic acid</p>	<p>2,4-dichlorobut-1-ene</p> $\text{H}_2\text{C}=\text{C}(\text{Cl})-\text{CH}_2-\text{CH}_2(\text{Cl})$ <p>alkene</p>



1.

Numbering of the chain starts from the end that carries the main functional group, the double bond. Once counted from this end, the number of the double bond and chlorine change.

2-chloro-3-methylpent-1-ene.